#### Chemistry Unit 5 Study Guide:

Significant Figure: each of the digits of a number that are used to express it to the required degree of accuracy. starting from nonzero digit.

#### Rules for determining sig figs:

- a) All non-zero numbers are significant
- b) Zeros between non-zero digits are significant ex 7003
- c) Training zeros (those at the end) are only significant if the number contains a decimal point. Ex: 1,000.0=5 sig figs 7,000=1 sig fig
- d) Zeros to the left of the first non-zero are insignificant, they serve as place holder. Ex: 00.01

#### Calculations with sig figs:

Multiplication/Division:

Your answer should have the same amount of sig figs as the measurement with the fewest significant figures.

Ex: 3.021m x 9.0000m x 8.2m = 222.9498m<sup>3</sup>→220 because only has 2 sig figs.

Addition/subtraction:

Your answer should have the same number of digits to the right of the decimal as the measurement having the fewest number of digits to the right decimal.

Ex:  $5.001m + 98.5m + 5001.2315 = 5104.7325 \rightarrow 5104.7$ 

a x 10 n -> Exponent or Power Scientific Notation; a way for scientists to handle a very large or very small number in a more organized manner. Needs a coefficient (1-9.99), base, and exponent.

Negative exp- small number- mover to the left

Positive exp- big number- move to the right

Moles: in chemistry they are 6.02x1022 particles (atoms, molecules, or formula units) of a substance (elements and compunds)

1 mole =  $6.02 \times 10^{23}$ 

Converting between molecules and particles:

- a) Figure out which units you are converting between
- b) Set up a proportion
- c) Look at what the problem gives you. Multiply this by a coinversion factor formed from your equality in step 2.

Ex: how many moles of Mg is 1.25x10<sup>23</sup> atoms of magnesium.

a) Atoms-moles

 $1/6.02 \times 10^{23} = x/1.25 \times 10^{23}$  x has to be the same amount of sig fig as the number in the problem.

Converting Between moles and mass:

Grand Formula Mass/Molar Mass: the mass of one mole of a substance. Expressed in grams.

- Calculate GFM.
- Create conversion factor:

# **Unit 5 Study Guide Chemistry Answers**

**Rachel S Tattersall** 

#### **Unit 5 Study Guide Chemistry Answers:**

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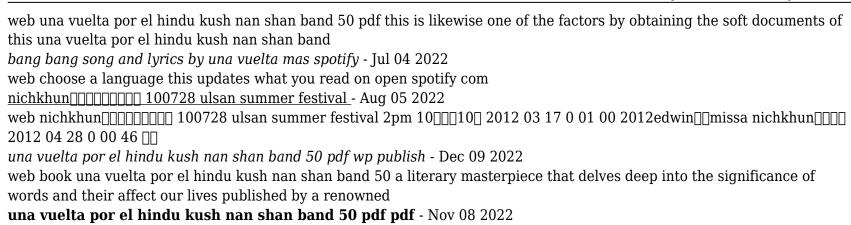
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