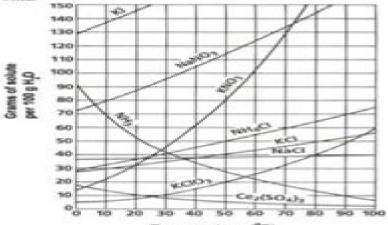
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Solubility Curve Practice Problems

Directions: Find the mass of solute will dissolve in 100mL of water at the following temperatures?

- KNO₃ at 70°C = ______
- NaCl at 100°C=______
- NH_iCl at 90°C=_____
- Which of the above three substances is most soluble in water at 15°C. =
- Which of the above three substances is lest soluble in water at 15°C. =



Period:

	a desirebasis de financia de f	
2. Solution	Saturated, super saturated or Unsaturated?	If unsaturated: How much more solute can dissolve in the solution?
a solution that contains 70g of NaNO3 at 30°C (in 100 mL H2O)		
a solution that contains 50g of NH ₄ Cl at 50°C (in 100 mL H ₂ O)		
a solution that contains 20g of KClO ₃ at 50°C (in 50 mL H ₂ O)		
a solution that contains 70g of KI at 0°C (in 200 mL H ₂ O)		

- 3. a. How many grams of KCl can dissolve at 5°C?
 - b. How many grams of KCl can dissolve at 25°C?
 - c. How many grams of Ce₂(SO₄)₃ can dissolve at 10°C?
 - d. How many grams of Ce₂(SO₄)₃ can dissolve at 50°C?
- 4. At 90°C, you dissolved 10 g of KCl in 100, g of water. Is this solution saturated or unsaturated?
- 5. A mass of 100 g of NaNO3 is dissolved in 100 g of water at 80°C.
 - a) Is the solution saturated or unsaturated?
 - b) As the solution is cooled, at what temperature should solid first appear in the solution? Explain.
- 6. Use the graph to answer the following two questions:

Which compound is most soluble at 20 °C?

Which is the least soluble at 40 °C?

Which substance on the graph is least soluble at 10°C?

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