

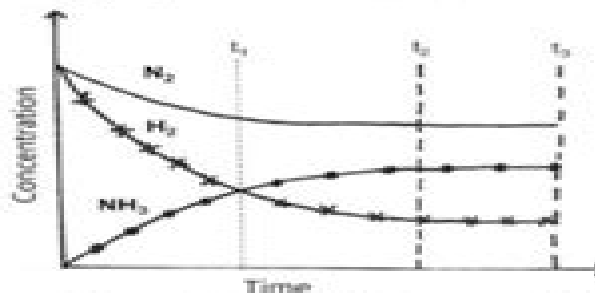
Chemical Equilibrium

The Equilibrium Condition

- The majority of chemical reactions are reversible.
- Equilibrium is the state where the **rate of the forward reaction is equal to the rate of the reverse reaction** and results in no net observable change. At equilibrium, concentrations of all reactants and products remain constant at constant temperature.
- Equilibrium is dynamic.
 - $\rightleftharpoons$ or $\rightleftharpoons$ indicates that the reaction is proceeding forward and in reverse.
 - Once equilibrium is established, the rate of each direction is equal.
 - This keeps the concentration of reactants constant
 - This keeps the concentration of products constant
 - This does NOT mean that concentration of products and concentration of reactants are the same or equal to each other.
- Any chemical reactions carried out in a closed vessel will reach equilibrium.
- Concept Check



H_2 gas and N_2 gas were placed in a rigid vessel and allowed to reach equilibrium in the presence of a catalyst according to the equation above. The diagram below shows how the concentrations of H_2 , N_2 , and NH_3 in this system changed over time. Based on the graph, which of these claims is most likely correct?



Temp
Dependent

- The reaction reached equilibrium at t_1 , because the concentrations of H_2 and NH_3 were the same.
- The reaction reached equilibrium at t_2 , because the concentrations of N_2 , H_2 , and NH_3 stopped changing.
- The reaction reached equilibrium at t_3 , because the reaction stopped proceeding.
- The reaction did not reach equilibrium, because not all of the N_2 and H_2 was consumed at time t_3 .

Unit 4 Chemical Equilibrium Worksheet

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