

Chapter 4, Section 4: Reaction types – Acid-Base reactions

EX 12: What volume of 0.100 M HCl solution is needed to neutralize 25.0 mL of 0.350 M NaOH? Then replace HCl with H_2SO_4 ?



EX 13: In an experiment, 28 mL of 0.250 M HNO_3 and 53 mL of 0.320 M KOH are mixed. Calculate the amount of water formed in the reaction and calculate the concentration of H^+ or OH^- in excess after the reaction goes to completion.

Acid-Base

$$(28 \text{ mL} \times 0.250 \text{ M}) (1) = (53 \text{ mL} \times 0.320 \text{ M}) (1)$$

Acid $\rightarrow$ $\frac{0.250 \text{ M} \times 0.028 \text{ L}}{1 \text{ L}} = 0.007 \text{ mol } HNO_3$

Base $\rightarrow$ $\frac{0.320 \text{ M} \times 0.053 \text{ L}}{1 \text{ L}} = 0.017 \text{ mol } KOH$

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