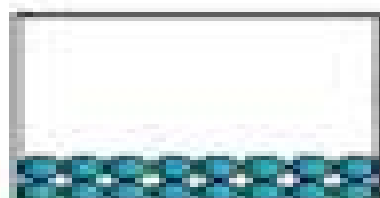
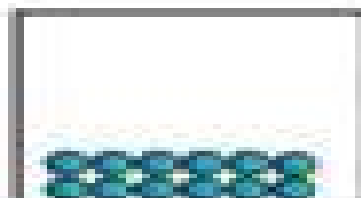


ENTROPY, DISORDER

Entropy is the degree of randomness in a substance. This means that change in entropy in ΔS , which are very ordered and have low entropy. Molecules with smaller bond lengths have stronger forces that they come about, more closely, and gases have an even larger amount of entropy. According to the second law of thermodynamics, nature always progresses in a state of highest entropy.



Entropy increases when the following reactions occur and decrease in enthalpy in different conditions ΔH or ΔG .



HEAT OF FORMATION, ENTHALPY

The equation for heat of formation is $\Delta H = \Delta G - T\Delta S$.

This is known as the spontaneity, the sign for ΔH that is the negative, ΔH represents the heat of reaction, ΔS is the change in entropy, T is temperature in Kelvin.

A negative value for ΔH means that the reaction is exothermic. This means that heat is released.

A positive value for ΔH means that the reaction is endothermic. This means that heat is absorbed.

A negative value for ΔS means that the products are more ordered than the reactants.

A positive value for ΔS means that the products are less ordered than the reactants.

Unit 15 Reaction Energy And Reaction Kinetics

M Carnoy



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Progress in Reaction Kinetics G. Porter, 2013-09-11 Progress in Reaction Kinetics Volume 3 presents articles about advances in reaction kinetics The book begins with a theoretical review of bimolecular reactions such as the relation between free energy and potential energy surfaces The text describes reactions of hydrogen atoms in the gas phase the hot atom chemistry of gas phase systems the inhibition of gaseous free radical chain reactions and vibrational relaxation in gases Articles about pulse radiolysis the effects of dose rate and linear energy transfer in radiation chemistry and the electronic spectra and kinetics of aromatic free radicals are also considered The book further tackles the kinetics of polymerization of vinyl monomers by lithium alkyls as well as radical polymerization in solutions Chemists and professionals dealing with radiation physical and industrial chemistry will find the book invaluable **Cambridge International AS and A Level**

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