

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES  
(CHEMISTRY)**

**TABLE 1      PHYSICAL CONSTANTS**

| NAME                            | SYMBOL    | VALUE                                     |
|---------------------------------|-----------|-------------------------------------------|
| Magnitude of charge on electron | $e$       | $1,6 \times 10^{-19} \text{ C}$           |
| Mass of an electron             | $m_e$     | $9,1 \times 10^{-31} \text{ kg}$          |
| Standard pressure               | $p^\circ$ | $1,01 \times 10^5 \text{ Pa}$             |
| Molar gas volume at STP         | $V_m$     | $22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$ |
| Standard temperature            | $T^\circ$ | $273 \text{ K}$                           |
| Avogadro's constant             | $N_A$     | $6,02 \times 10^{23} \text{ mol}^{-1}$    |
| Faraday's constant              | $F$       | $96\,500 \text{ C} \cdot \text{mol}^{-1}$ |

**TABLE 2      CHEMISTRY FORMULAE**

|                                         |                                                                                                                                                                             |                     |
|-----------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------------------|
| $n = \frac{m}{M}$                       | $n = \frac{N}{N_A}$                                                                                                                                                         | $n = \frac{V}{V_m}$ |
| $c = \frac{n}{V}$ OR $c = \frac{m}{MV}$ | $K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$<br>at 25 °C (298 K)                                                                                  |                     |
| $q = It$<br>$q = nF$                    | $E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$<br>$E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$ |                     |

# Tmrf Exemplar Physical Science P2

**Paul G. Hewitt, John A. Suchocki, Leslie  
A. Hewitt**



## **Tmrf Exemplar Physical Science P2:**

**Loose Leaf Physical Science** Bill Tillery, 2011-01-13 Physical Science Ninth Edition is a straightforward easy to read but substantial introduction to the fundamental behavior of matter and energy It is intended to serve the needs of non science majors who are required to complete one or more physical science courses It offers exceptional straight forward writing complemented with useful pedagogical tools Physical Science introduces basic concepts and key ideas while providing opportunities for students to learn reasoning skills and a new way of thinking about their environment No prior work in science is assumed The text offers students complete coverage of the physical sciences with a level of explanation and detail appropriate for all students The sequence of chapters in Physical Science is flexible and the instructor can determine topic sequence and depth of coverage as needed The materials are also designed to support a conceptual approach or a combined conceptual and problem solving approach Along with the accompanying laboratory manual the text contains enough material for the instructor to select a sequence for a two semester course It can also serve as a text in a one semester physics and chemistry course      Conceptual Physical Science Masteringphysics Standalone Access Card Paul G. Hewitt, John A. Suchocki, Leslie A. Hewitt, 2011-12-01      **Introduction to Concepts and Theories in Physical Science** Gerald Holton, 1985

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