

1. Prepping the crucibles and crucible lids. Observe, clean and dry two crucibles and their lids. As mentioned in the safety precautions, the crucibles will get hot when heated and feel changes color to marooned in the safety precautions, the crucibles with their lids on *you* use clay triangles. The lid should be opened hot. Using tongs, place the crucibles with their lids on *you* use clay triangles. Then heat the crucibles with their lids slightly to allow water vapor to easily escape the crucible. Then heat the crucibles with their lids off for about four minutes. The point of this is to make sure the crucibles and their lids are completely dry. Turn the burner off and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Using tongs, carry the crucibles to your bench. **THE CRUCIBLES ARE STILL HOT!** Wait 10 minutes while the crucibles and lids further cool. While the crucibles and lids are cooling proceed to step 2.
2. Obtain 1.0 grams of each hydrate and place each in a separate crucible, after the crucibles have cooled. $\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$ is the blue sample, while $\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$ is the purple sample. Put the crucible lids on top of the crucibles. Weigh each crucible, its lid and the hydrate, record these values in the data table at the top of page 3. In each sample x is an unknown integer.
3. Place each crucible with its lid on a clay support triangle and heat each for five minutes on low flame. After this slightly increase the flame and continue heating for ten minutes. After this time, turn off the burner and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Using tongs, carry the crucible burner and wait 10 minutes for the crucible to cool, increasing the temperature slowly is to minimize the probability of spattering of the sample during dehydration. The lid is to prevent the loss of sample in case of some spattering. The cobalt chloride hydrate often does spatter if heated rapidly. Lastly, do not let the crucible get red hot since unwanted sample decomposition may occur.
4. Carry the crucible to the balance with tongs, the crucibles still will be slightly hot. Weigh each crucible lid and now anhydrous sample. Record these values.
5. Repeat steps three and four. Our goal here is to drive off any remaining water in the sample. If the difference in the masses between the two measurements is greater than 0.01 g, steps three and four must be repeated yet again. Record these values and all values in Table 1.
6. Keep a note of the colors of the samples before and after dehydration.

TABLE 1
DATA TABLE

Parameter	$\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$	$\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$
Mass of crucible and lid (g)	N/A	N/A
Mass of crucible, lid and sample (g)	N/A	N/A
Mass of sample (g)	1.000 g	1.000 g
Mass of crucible, lid and sample after first heating (g)	No Entry Here	No Entry Here
after second heating (g)	0.718 g	0.731 g
Mass of water lost (g)	0.282 g	0.269 g
Percent of water in sample by mass	28.2%	26.9%

Way To Go Indigo Lab Pre Lab Answers

Debbie Hennessey, Miranda Taylor



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