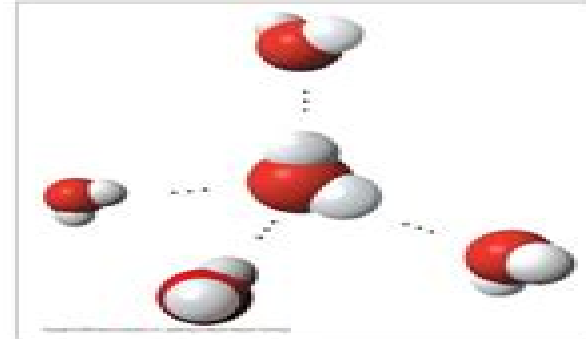


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Chapter 3: Water and Life**Concept 3.1** *The polarity of water molecules results in hydrogen bonding*

1. Study the water molecules at the right. On the central molecule, label oxygen (O) and hydrogen (H).
2. What is a *polar molecule*? Why is water considered polar?
3. Now, add + and – signs to indicate the charged regions of *each* molecule. Then, indicate the hydrogen bonds.
4. Explain *hydrogen bonding*. How many hydrogen bonds can a single water molecule form?

**Concept 3.2** *Four emergent properties of water contribute to Earth's fitness for life*

Hydrogen bonding accounts for the unique properties of water. Let's look at several.

Cohesion

5. Distinguish between *cohesion* and *adhesion*.
6. What is demonstrated when you see beads of water on a waxed car hood?
7. Which property explains the ability of a water strider to walk on water?

Moderation of Temperature

8. The calorie is a unit of heat. Define *calorie*.
9. Water has high *specific heat*. What does this mean? How does water's specific heat compare to alcohol's?
10. Explain how hydrogen bonding contributes to water's high specific heat.

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