

Unit 6: Thermodynamics Study Guide and Practice Test

Topics for this exam:

- 6.1 Endothermic and Exothermic Processes
- 6.2 Energy Diagrams
- 6.3 Heat Transfer and Thermal equilibrium
- 6.4 Heat Capacity and Calorimetry
- 6.5 Energy of Phase Changes
- 6.6 Enthalpy of Reaction
- 6.7 Bond Enthalpy
- 6.8 Enthalpy of Formation
- 6.9 Hess's Law

Vocabulary

- | | | |
|-----------------------|--------------------|-------------------------|
| • System | • Endothermic | • Calorie |
| • Surroundings | • Exothermic | • Standard State |
| • Universe | • Kinetic Energy | • Hess's law |
| • State Function | • Potential Energy | • Enthalpy of Formation |
| • Energy | • Specific Heat | • Enthalpy of Reaction |
| • Enthalpy | • Heat Capacity | |
| • Entropy | • Work | |
| • Gibbs's Free Energy | • Joule | |

Practice Test- Multiple Choice

1) Which of the following reactions have a negative entropy change?

- I. $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l})$
II. $2 \text{NH}_3(\text{g}) \rightarrow \text{N}_2(\text{g}) + 3 \text{H}_2(\text{g})$
III. $\text{Ca}(\text{s}) + \text{Cl}_2(\text{g}) \rightarrow \text{CaCl}_2(\text{s})$

- (A) II only
(B) I only

- (C) I, II, and III
(D) I and III

2) A certain reaction is non-spontaneous under standard conditions, but becomes spontaneous at higher temperatures. What conclusions may be drawn under standard conditions?

- (A) $\Delta H < 0$, $\Delta S > 0$, and $\Delta G > 0$
(B) $\Delta H > 0$, $\Delta S < 0$, and $\Delta G > 0$

- (C) $\Delta H > 0$, $\Delta S > 0$, and $\Delta G > 0$
(D) $\Delta H < 0$, $\Delta S < 0$, and $\Delta G > 0$

3) The value of ΔH_{rxn} for the reaction below is -126 kJ. How much heat (in kJ) is released when 2.00 mol of NaOH is formed in the reaction?



- A) 252 B) -126 C) 7.8 D) 63 E) 3.9

4) $3 \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_6\text{H}_6(\text{g})$
What is the standard enthalpy change, ΔH° , for the reaction represented above?
(ΔH_f° of $\text{C}_2\text{H}_2(\text{g})$ is 230 kJ mol⁻¹; ΔH_f° of $\text{C}_6\text{H}_6(\text{g})$ is 83 kJ mol⁻¹.)

- (A) -607 kJ (B) -147 kJ (C) -19 kJ (D) +19 kJ

5) $\text{CH}_4(\text{g}) + 2 \text{O}_2(\text{g}) \rightarrow \text{CO}(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$; $\Delta H_{\text{rxn}}^\circ = - 889.1 \text{ kJ}$
 $\Delta H_f^\circ \text{H}_2\text{O}(\text{l}) = -285.8 \text{ kJ / mole}$

Thermodynamics Study Guide Answers

Dilip K. Das, Rajaram K. Prabhudesai



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Study Guide/Solutions Manual for Organic Chemistry K. Peter C. Vollhardt, Neil E. Schore, 2018-01-12 Updated for the Eighth Edition of Vollhardt Schore Organic Chemistry and written by the book's coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

Thermodynamics Ibrahim Dincer, 2020-08-18 Presents a unique stepwise exergy based approach to thermodynamic concepts systems and applications Thermodynamics A Smart Approach redefines this crucial branch of engineering as the science of energy and exergy rather than the science of energy and entropy to provide an innovative step by step approach for teaching understanding and practicing thermodynamics in a clearer and easier way Focusing primarily on the concepts and balance equations this innovative textbook covers exergy under the second law of thermodynamics discusses exergy matters and relates thermodynamics to environmental impact and sustainable development in a clear simple and understandable manner It aims to change the way thermodynamics is taught and practiced and help overcome the fear of thermodynamics Author Ibrahim Dincer a pioneer in the areas of thermodynamics and sustainable energy technologies draws upon his multiple decades of experience teaching and researching thermodynamics to offer a unique exergy based approach to the subject Enabling readers to easily comprehend and apply thermodynamic principles the text organizes thermodynamics into seven critical steps property state process cycle first law of thermodynamics second law of thermodynamics and performance assessment and provides extended teaching tools for systems and applications Precise student friendly chapters cover fundamental concepts thermodynamic laws conventional and innovative power and refrigeration cycles and more This textbook Covers a unique approach in teaching design analysis and assessment of thermodynamic systems Provides lots of examples for every subject for students and instructors Contains hundreds of illustrations figures and tables to better illustrate contents Includes many conceptual questions and study problems Features numerous systems related examples and practical applications Thermodynamics A Smart Approach is an ideal textbook for undergraduate students and graduate students of engineering and applied science as well researchers scientists and practicing engineers seeking a precise and concise textbook and or reference work **Fundamentals of Chemistry, Study Guide** James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science

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