

Chapter 4, Section 4: Reaction types – Acid-Base reactions

EX 12: What volume of 0.100 M HCl solution is needed to neutralize 25.0 mL of 0.350 M NaOH? Then replace HCl with H_2SO_4 ?



EX 13: In an experiment, 28 mL of 0.250 M HNO_3 and 53 mL of 0.320 M KOH are mixed. Calculate the amount of water formed in the reaction and calculate the concentration of H^+ or OH^- in excess after the reaction goes to completion.

Acid-Base

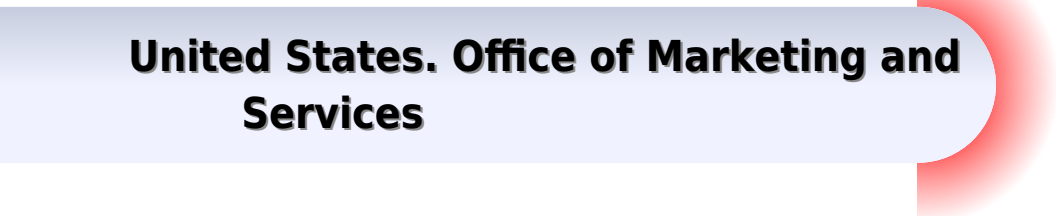
$$(28 \text{ mL} \times 0.250 \text{ M}) (1) = (53 \text{ mL} \times 0.320 \text{ M}) (1)$$

Acid $\rightarrow$ $\frac{0.250 \text{ M} \times 0.028 \text{ L}}{1 \text{ L}} = 0.007 \text{ mol } HNO_3$

Base $\rightarrow$ $\frac{0.320 \text{ M} \times 0.053 \text{ L}}{1 \text{ L}} = 0.017 \text{ mol } KOH$

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Table of Contents Unit 4 Section A Supplement Answers Chemistry

1. Understanding the eBook Unit 4 Section A Supplement Answers Chemistry
 - The Rise of Digital Reading Unit 4 Section A Supplement Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Section A Supplement Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 4 Section A Supplement Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 4 Section A Supplement Answers Chemistry
 - Personalized Recommendations
 - Unit 4 Section A Supplement Answers Chemistry User Reviews and Ratings
 - Unit 4 Section A Supplement Answers Chemistry and Bestseller Lists
5. Accessing Unit 4 Section A Supplement Answers Chemistry Free and Paid eBooks
 - Unit 4 Section A Supplement Answers Chemistry Public Domain eBooks
 - Unit 4 Section A Supplement Answers Chemistry eBook Subscription Services
 - Unit 4 Section A Supplement Answers Chemistry Budget-Friendly Options
6. Navigating Unit 4 Section A Supplement Answers Chemistry eBook Formats

- ePub, PDF, MOBI, and More
 - Unit 4 Section A Supplement Answers Chemistry Compatibility with Devices
 - Unit 4 Section A Supplement Answers Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Unit 4 Section A Supplement Answers Chemistry
 - Highlighting and Note-Taking Unit 4 Section A Supplement Answers Chemistry
 - Interactive Elements Unit 4 Section A Supplement Answers Chemistry
8. Staying Engaged with Unit 4 Section A Supplement Answers Chemistry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 4 Section A Supplement Answers Chemistry
9. Balancing eBooks and Physical Books Unit 4 Section A Supplement Answers Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 4 Section A Supplement Answers Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 4 Section A Supplement Answers Chemistry
- Setting Reading Goals Unit 4 Section A Supplement Answers Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 4 Section A Supplement Answers Chemistry
- Fact-Checking eBook Content of Unit 4 Section A Supplement Answers Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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