

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

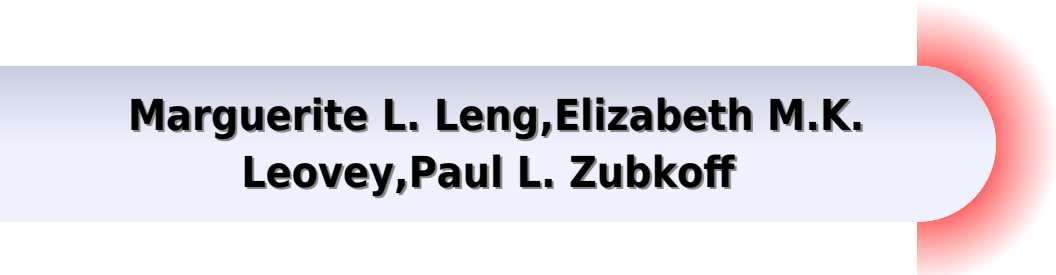
- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

Unit 6 Section D Content Guide Chemistry

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Unit 6 Section D Content Guide Chemistry:

Edexcel AS/A2 Chemistry Student Unit Guide: Units 3 and 6 Chemistry Laboratory Skills George Facer, 2010-02-26 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements This guide offers advice on preparing for the Edexcel Unit 3 and Unit 6 Chemistry Laboratory Skills assessments Agrochemical Environmental Fate State of the Art Marguerite L. Leng, Elizabeth M.K. Leovey, Paul L. Zubkoff, 1995-04-28 Accurate risk assessment is critical to pesticide regulation This authoritative reference provides an exhaustive evaluation of current agrochemical environmental fate studies a critical review of current EPA pesticide assessment guidelines and a wide variety of environmental simulation models Divided into four sections this well organized book provides a wealth of data and information vital to anyone involved in environmental exposure assessment groundwater surface water and water contamination pesticide regulation and environmental simulation modeling At your fingertips you will have the latest information on the development of meaningful environmental fate data and how this information will result in accurate assessment of potential environmental and human hazards The inadequacy of current regulatory guidelines and the resulting nonscientific assessment of agrochemical environmental fate are discussed in detail A wide variety of environmental fate studies are included to demonstrate the current use of data to assess environmental fate and potential hazards associated with agrochemical use Finally ten chapters discuss the use of computer models that have been developed for analyzing and integrating data from a variety of environmental fate studies on agrochemicals used under various field conditions **Resources in Education** ,1995 Fulton Generating Station Patsy L. Hendricks, 1974 **Educational Research Document Summaries** Educational Research Information Center (U.S.), 1966 **Research in Education** ,1974 **Industrial & Engineering Chemistry** ,1926 Chemical Engineering Catalog ,1918 *Clinton Power Station* Faye H. Horn, 1974 *Code of Federal Regulations* ,1996

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