

AP Chemistry: Kinetics Practice Problems

Directions: Write your answers to the following questions in the space provided. For problem solving, show all of your work. Make sure that your answers show proper units, notation, and significant digits. Do not use a calculator on the multiple choice questions.

- Three major methods used to increase the rate of a reaction are adding a catalyst, increasing the temperature, and increasing the concentration of a reactant. From the perspective of collision theory, explain how each of these methods increases the reaction rate.
 - adding a catalyst
Adding a suitable catalyst lowers the energy demands of the reaction (the activation energy) by providing a less demanding pathway. It does this by forming a different activated complex with a lower activation energy.
 - increasing the temperature
Increasing the temperature increases the average kinetic energy of the molecules. However, what is most important here is that the fraction of molecules that have the energy necessary to react (the activation energy) is increased.
 - increasing the concentration of the reactants
Increasing the concentration of a reactant increases the probability of collisions, and therefore increases the possibility of reaction, for the particles must collide to react. Note that not all collisions result in successful reactions.
- Why do large crystals of sugar burn more slowly than finely ground sugar?
The rate of combustion is proportional to the surface area of sugar exposed to oxygen. Smaller crystals have more surface area and burn faster.
- How do homogeneous catalysts and heterogeneous catalyst differ?
A homogeneous catalyst is a catalyst that exists in the same phase (liquid or gas) as the reactants. A heterogeneous catalyst, on the other hand, is a catalyst that exists in a different phase from the reactants.
- Express the rate of reaction in terms of the rate of change of each reactant and each product in the following.
 - $3\text{ClO}^-(\text{aq}) \rightarrow \text{ClO}_3^-(\text{aq}) + 2\text{Cl}^-(\text{aq})$
$$\text{rate of reaction} = -\frac{\Delta[\text{ClO}^-]}{3\Delta t} = \frac{\Delta[\text{ClO}_3^-]}{\Delta t} = \frac{\Delta[\text{Cl}^-]}{2\Delta t}$$
 - $3\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
$$\text{rate of reaction} = -\frac{\Delta[\text{SO}_2]}{3\Delta t} = -\frac{\Delta[\text{O}_2]}{\Delta t} = \frac{\Delta[\text{SO}_3]}{2\Delta t}$$
 - $\text{C}_2\text{H}_4(\text{g}) + \text{Br}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4\text{Br}_2(\text{g})$
$$\text{rate of reaction} = -\frac{\Delta[\text{C}_2\text{H}_4]}{\Delta t} = -\frac{\Delta[\text{Br}_2]}{\Delta t} = \frac{\Delta[\text{C}_2\text{H}_4\text{Br}_2]}{\Delta t}$$
- In the Haber process for the production of ammonia,
$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$$

What is the relationship between the rate of production of ammonia and the rate of consumption of hydrogen?
$$\text{rate of reaction} = -\frac{\Delta[\text{H}_2]}{3\Delta t} = \frac{\Delta[\text{NH}_3]}{2\Delta t}$$
- The rate constant
 - always shows an exponential increase with the Kelvin or absolute temperature.
 - increases with increasing concentration.
 - usually increases with increased pressure for gases.
 - never changes (it is a constant).
 - is the same for a given reaction at the same Kelvin or absolute temperature.

Answer: (E) In this question be sure to note that it is the rate constant that is involved. Refer to the Arrhenius equation for the specific relationship.

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L Reisser



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Table of Contents Unit 3 Petroleum Study Guide Answers

1. Understanding the eBook Unit 3 Petroleum Study Guide Answers
 - The Rise of Digital Reading Unit 3 Petroleum Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 3 Petroleum Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 3 Petroleum Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 3 Petroleum Study Guide Answers
 - Personalized Recommendations

- Unit 3 Petroleum Study Guide Answers User Reviews and Ratings
- Unit 3 Petroleum Study Guide Answers and Bestseller Lists
- 5. Accessing Unit 3 Petroleum Study Guide Answers Free and Paid eBooks
 - Unit 3 Petroleum Study Guide Answers Public Domain eBooks
 - Unit 3 Petroleum Study Guide Answers eBook Subscription Services
 - Unit 3 Petroleum Study Guide Answers Budget-Friendly Options
- 6. Navigating Unit 3 Petroleum Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 3 Petroleum Study Guide Answers Compatibility with Devices
 - Unit 3 Petroleum Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 3 Petroleum Study Guide Answers
 - Highlighting and Note-Taking Unit 3 Petroleum Study Guide Answers
 - Interactive Elements Unit 3 Petroleum Study Guide Answers
- 8. Staying Engaged with Unit 3 Petroleum Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 3 Petroleum Study Guide Answers
- 9. Balancing eBooks and Physical Books Unit 3 Petroleum Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 3 Petroleum Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 3 Petroleum Study Guide Answers
 - Setting Reading Goals Unit 3 Petroleum Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 3 Petroleum Study Guide Answers
 - Fact-Checking eBook Content of Unit 3 Petroleum Study Guide Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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