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Marcel A. Müller



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web feb 24 2019 question 1 it is important to be able to recognize the major types of chemical reactions comstock getty images the chemical reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ is a a synthesis reaction b decomposition reaction c single displacement reaction d double displacement reaction e combustion reaction question 2

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web an answer key document has been included and is available for download some additional information to be aware of for teachers demonstrations 1 and 5 will not produce a chemical reaction five of the reactions will produce indicators of a chemical reaction for students to observe

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web identifying chemical reaction types key part 1 identify the type of reaction 1 $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$ double displacement 2 $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$ double displacement 3 $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ combustion 4 $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ single displacement 5 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ decomposition 6 $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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web chemical equations can also be used to represent physical processes write a chemical reaction for the boiling of water including the proper phase labels chemical equations can also be used to represent physical processes write a chemical reaction for the freezing of water including the proper phase labels explain why $4\text{Na} + \text{S} \rightarrow 2\text{Na}_2\text{S}$

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web 1 $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement 2 $3\text{Ca}(\text{OH})_2 + \text{Al}_2(\text{SO}_4)_3 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$ type of reaction double displacement 3 $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$ type of reaction single displacement 4 $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ type of reaction combustion 5 $2\text{PbSO}_4 + 2\text{PbSO}_3 + \text{O}_2$ type of reaction

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