_22.5 L O₂ ____10. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of Dinitrogen trioxide.

N_2	+	302	 2N20
		+	+
		most	 mol

15.0 L N203	1 mol-N ₂ O ₃	3 mgl-05	22.4 L O ₂
11	22.4 L-N2O3	2 mol-N ₂ O ₃	1 mol-O₂

Write the correct answer for each of the following statements or questions in the blank provided. Use your book, if necessary, to find the terms that fit the definitions.

6.02 X 10 ²³	1.1.	Avogadro's number?
-------------------------	------	--------------------

Reactants	15.	the starting materials in an equation.
Mass-mass		
Calculations	16.	determining the mass of products from the given mass of
	100.00	ta rets

Multiple Choice

- B 17. Which of these expressions is an incorrect interpretation of the balanced equation? 2S + 3O₂ → 2SO₃
 - A. 2 atoms S + 3 molecules O₂ → 2 molecules S O₃
 - B. $2gS + 3gO_2 \rightarrow 2gSO_3$
 - C. 2 mole S + 3 mole O₂ → 2 mole SO₃
 - D. none of the above
- _C___18. In a chemical reaction, the mass of the products:
 - is less than the mass of the reactants.
 - Is greater than the mass of the reactants.
 - Is equal to the mass of the reactants.
 - D. Has no relationship to the mass of the reactants.

Unit 8 Stoichiometry Answers

Yan Bai

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